

CHEMISTRY WORKSHEET

CLASS: 8

1. Name the following:

- a. A soft metal
- b. Chemical name of baking soda
- c. An inert gas with four shells
- d. A liquid non metal
- e. A solid-solid homogeneous mixture
- f. A liquid-liquid homogeneous mixture
- g. The scientist who gave plum-pudding model of atom
- h. The amount of charge on an electron in coulombs
- i. The property of an atom denoted by 'Z'
- j. The rule that denotes maximum number of electrons in a shell
- k. The number of electrons given or taken by an atom of an element during reaction
- l. Symbol of phosphate radical
- m. A chemical equation with only one molecule of reactants and products
- n. A catalyst that decreases the rate of reaction
- o. An indicator that turns red in acidic and yellow in basic medium
- p. An amphoteric oxide
- q. An example of non-metallic alkali
- r. An example of an anion having metal atom
- s. The scientist who called hydrogen "The inflammable gas"
- t. The electrode connected to the positive terminal of battery
- u. A combustible gas that is a non-supporter of combustion
- v. A substance that supplies hydrogen or removes oxygen
- w. A chemical process that involves loss of electrons
- x. Composition of water gas
- y. Liquefaction temperature and pressure of hydrogen
- z. A chemical reaction between two compounds by exchange of their radicals

2. Define/Short notes on the following with an example of each:

- a. Displacement reaction
- b. Promoter
- c. Endothermic reaction
- d. Effervescence
- e. Variable valency
- f. Rutherford's atomic model
- g. Fractional distillation
- h. Crystallization
- i. Water of crystallization
- j. Anomalous expansion of water

3. Give reasons:

- a. Atom is stable despite positively and negatively charged particles.
- b. Isotopes have similar chemical property.
- c. Metals and non-metals react but noble gases do not react.
- d. It is necessary to balance a chemical reaction.
- e. Elements from a compound cannot be separated by physical methods.
- f. Wasp sting is treated with vinegar.
- g. Toothpastes are basic in nature.
- h. All bases are not alkalis.
- i. Nitric acid is not used for preparation of hydrogen.
- j. Silica gel is kept in unused water bottles.
- k. Hard water is not fit for laundries.
- l. Ice floats on water.
- m. Water pipes burst in severe winters.

4. Differentiate with an example:

- a. Miscible and immiscible liquids
- b. Saturated and unsaturated solution
- c. Anhydrous and hydrated substances
- d. Efflorescent and hygroscopic substances
- e. Hard and soft water

5 a. Write balanced chemical equation for each of the following reactions:

- i. Laboratory preparation of CO_2
- ii. Heating of CaCO_3
- iii. Action of dil. Sulphuric acid on sodium carbonate
- iv. Carbon dioxide with burning magnesium ribbon
- v. Reaction of CO_2 with sodium oxide
- vi. Dissolution of CO_2 in water
- vii. Heating of sodium bicarbonate
- viii. Reaction of caustic potash and CO_2
- ix. Burning of CO
- x. Reaction of CO and ferric oxide

5b. Define:

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| i. Allotropes | vii. carbogen |
| ii. Catenation | viii. Hopcolite |
| iii. Graphite | ix. Carbonisation |
| iv. Fullerene | x. Adsorption |
| v. Anthracite | xi. Sugar charcoal |
| vi. Adsorption | xii. Carat |

